



## An efficient catalytic method for fulvene synthesis

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### ABSTRACT

The effects of the nature and amount of base, substrate structure, amount of added water and solvent on the condensation of carbonyl compounds with cyclopentadiene in the presence of secondary amines were investigated. Based on these studies, a new efficient and green synthesis of fulvenes was developed.

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### 1. Introduction

Fulvenes constitute an important class of cross-conjugated olefins with unique electronic, spectroscopic and chemical properties.<sup>1</sup> Though the first fulvene synthesis was reported by Thiele in 1900,<sup>2</sup> considerable effort has been directed toward synthetic improvements,<sup>3,4</sup> since Thiele's original method based on the condensation of aldehydes and ketones with cyclopentadiene in the presence of alkoxides gives low yields with considerable amounts of resinous byproducts, presumably due to competing aldol condensations.

Neuenschwander developed an alternative synthesis of fulvenes that are not readily accessible by the Thiele procedure.<sup>5,6</sup> Alkylation of sodium cyclopentadienide with 1-chloroalkyl acetates followed by elimination of acetic acid from acetoxyalkyl cyclopentadienes with base gives fulvenes in moderate to low yields. However, this method allows the synthesis and isolation of volatile fulvenes, such as the unsubstituted parent fulvene, 6-vinylfulvene and 6-ethynylfulvene since no aqueous work-up is necessary. This three-step method requires the synthesis of 1-chloroalkyl acetates and is also less suited for the synthesis of 6,6-dialkyl or diarylfulvenes.

Hafner's synthesis<sup>7</sup> of 6-monosubstituted fulvenes is based on alkyl or aryl lithium reagents onto 6-dimethylaminofulvene; since the pyrrolidine variant (Little and Stone, *vide infra*) with lower MW aldehydes, such as ethanal and propanal is unsatisfactory, Hafner's

two-step method appears to be the method of choice for the synthesis of 6-methylfulvene, a substrate most frequently used in organometallic chemistry for fulvene-based metal complexes as well as materials chemistry.<sup>8–12</sup> In this full report we also describe an exceptionally simple and practical synthesis of the aforementioned 6-methyl and 6-ethylfulvenes that allows their isolation in pure form without extensive work-up and chromatography (separation of these volatile fulvenes from chromatography solvents without significant losses is extremely difficult).

A method for the synthesis of volatile and sensitive fulvenes, such as the parent fulvene, 6-methyl and 6-ethylfulvene employing basic ion exchange resin has been reported, however, the authors were not able to isolate the aforementioned fulvenes from the reaction mixtures and their formation was inferred from the UV absorption of possible fulvene contents.<sup>13,14</sup>

Stone and Little reported a procedure whereby excess pyrrolidine was used as base with methanol as solvent and achieved good yields in the case of aliphatic aldehydes and ketones and aromatic aldehydes.<sup>15</sup> The Stone and Little variant has been widely used in the last 25 years (over 275 citations), and deservedly attained the status as 'the method of choice for fulvene synthesis' from simple aldehydes and non-aromatic ketones. Yet it has some limitations mostly brought on by the fact that (a) at least 1.5 equiv of the highly toxic and malodorous pyrrolidine and excess cyclopentadiene need to be used. Some of the other limitations include (b) conjugate addition of pyrrolidine onto the enone when  $\alpha,\beta$ -unsaturated ketones or aldehydes are used although this side reaction may be overcome in some instances by using diethylamine<sup>16</sup> (c) sterically hindered ketones, in particular aromatic ketones cannot be used; (d) sterically hindered

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enones as well as  $\beta$ -aryl enones undergo Michael attack by the cyclopentadienide ion.<sup>17</sup>

The only method utilizing catalytic amount of amine base was reported by Freiesleben in 1963.<sup>3</sup> The author claimed having achieved near quantitative yields with catalytic HNEt<sub>2</sub> (or any other base, such as methyl- and phenylhydrazine) without solvent. Among the systems that gave exceptionally high yields were also aromatic ketones,  $\alpha,\beta$ -unsaturated carbonyl compounds, including methyl vinyl ketone. In the two reports by Freiesleben, there was no spectral characterization of the products. These claims were later refuted by a careful study by Neuenschwander et al.<sup>5</sup> who found that the best yield that was achieved by the Swiss group under Freiesleben's conditions was 25% for 6,6-dimethylfulvene (instead of 99.8%), while in other cases the yields were 0–10%. Even methyl vinyl ketone was reported by Freiesleben in a patent to give the corresponding fulvene, however, Griesbeck found that this reaction leads to a Diels–Alder adduct rather than the fulvene.<sup>17</sup>

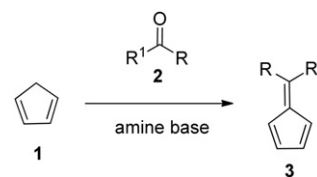
In 1995 we reported a survey of fulvene syntheses and implemented existing condensation methods (either Thiele or Little and Stone) in the syntheses of functionalized fulvenes as well as several 6-vinylfulvenes and explored the scope and limitations of each method.<sup>18</sup> More recently we disclosed a convenient and high yield synthesis of 6-methyl-6-vinylfulvene by a three-step protocol in an overall yield of 67% (previously available in only 0.5% yield by the Thiele procedure).<sup>19</sup>

In the course of our studies on the synthesis of new functionalized fulvenes, their implementation in organic synthesis and the chemistry of the endoperoxide derived from fulvenes<sup>20–24</sup> we needed an improved and environmentally friendly method for fulvene synthesis. Herein we report new efficient procedures for the preparation of a wide variety of fulvenes using catalytic pyrrolidine in MeOH/H<sub>2</sub>O that not only give high yields but also minimize product losses during extractive work-up and avoid the use of excess cyclopentadiene whose dimer presents challenging separation problems from the desired fulvenes. Under our conditions, even low molecular-weight fulvenes, such as 6-methyl and 6-ethylfulvene are prepared and isolated in high yields and for which the Little procedure proves unsatisfactory. In this report the scope and limitations of our catalytic variant are also discussed and alternative routes suggested. The optimization of the method is based on some mechanistic studies involving effect of base, solvent and substituent effects.

## 2. Results and discussion

### 2.1. Effect of base

The use of cyclopentadiene in excess causes the formation of dicyclopentadiene, which is troublesome in the purification process as well. Moreover, we found that the use of excess pyrrolidine prevents the crystallization of solid fulvenes in the aromatic aldehyde series. Thus, we decided to test whether less than stoichiometric amounts of pyrrolidine might be enough for the formation of the iminium ion intermediate and the use of NEt<sub>3</sub> would help the condensation/dehydration step in acceptable yields and rates. The reaction of **1a** was first conducted only in the presence of triethylamine, however no fulvene formation was observed after 3 h stirring at room temperature. Pyrrolidine (10 mol %) was then added and the reaction was followed by <sup>1</sup>H NMR; after 1 h the conversion was only 15%. However, performing the reaction in the presence of pyrrolidine and molecular sieves accelerated the reaction significantly (47% after 1 h). The reaction was quantitative after stirring overnight at room temperature. Moreover, the isolation of the product consisted of only vacuum filtration of the crude mixture and evaporation of the solvent to give pure fulvene (Scheme 1).



Scheme 1. Synthesis of fulvene derivatives.

A series of aromatic aldehydes (Table 1, entries a–j, method A) were subjected to condensation with cyclopentadiene at room temperature under the conditions mentioned above and the initial rates were determined at the 0.83rd h. The correlation of the initial rates ( $R_x$ ) with Hammett substituent constants  $\sigma$  provided a linear plot with  $\rho = -2.19$  (Fig. 1). As the plot clearly indicates, electron-donating groups accelerate the reaction while electron-withdrawing substituents decelerate it. In the cases of strong electron withdrawing groups, such as *p*-, and *m*-NO<sub>2</sub> groups the reactions do not proceed at all. The observed reactivity order of the aldehyde series was surprising since results from a previous study revealed that the reactivity order with nitrogen nucleophiles in THF is opposite to the one observed here.<sup>25</sup> Therefore, we followed the equilibria for hemiacetal formations of aromatic aldehydes **2a,b,g–i** in MeOD at room temperature in order to understand this reversal of reactivity (Table 3). The data revealed that in the cases of aromatic aldehydes with strong electron-withdrawing groups hemiacetal formation equilibria are shifted toward right, which

Table 1

Method development study on effects of NEt<sub>3</sub> and molecular sieves on pyrrolidine-catalyzed fulvene formation rates

Entry	R	R <sup>1</sup>	Method A: <sup>a</sup> % 0.83 h (after completion)	Method B: <sup>b</sup> % 0.83 h (after completion)
a	H	4-MeC <sub>6</sub> H <sub>4</sub>	47(98)	71(93)
b	H	4-MeOC <sub>6</sub> H <sub>4</sub>	70(97)	99(94)
c	H	4-Me <sub>2</sub> NC <sub>6</sub> H <sub>4</sub>	50(100)	67(97)
d	H	4-HOC <sub>6</sub> H <sub>4</sub>	9(85)	<sup>f</sup>
e	H	3-MeOC <sub>6</sub> H <sub>4</sub>	<sup>f</sup>	35(96)
f	H	3,4-(MeO) <sub>2</sub> C <sub>6</sub> H <sub>3</sub>	47(98)	79(99)
g	H	C <sub>6</sub> H <sub>5</sub>	23(96)	50(95)
h	H	4-ClC <sub>6</sub> H <sub>4</sub>	6(96)	18(96)
i	H	4-NO <sub>2</sub> C <sub>6</sub> H <sub>4</sub>	0 <sup>25</sup>	39 <sup>c</sup> (30) <sup>h</sup>
j	H	3-NO <sub>2</sub> C <sub>6</sub> H <sub>4</sub>	0	45 <sup>c</sup> (31)
k	Me	Me	69	86 (89)
l	Me	4-MeC <sub>6</sub> H <sub>4</sub>	0 <sup>d</sup>	(82) <sup>e</sup>
m	Ph	Ph	0 <sup>d</sup>	0
n	H	<i>i</i> -Propyl	35	64(88)
o	R=R <sup>1</sup> =cyclobutyl		55	80(90)
p	R=R <sup>1</sup> =cyclopentyl		51	77(85)
t	Me	(CH <sub>2</sub> ) <sub>2</sub> CO <sub>2</sub> H	74 <sup>g</sup>	
u	Me	(CH <sub>2</sub> ) <sub>3</sub> CO <sub>2</sub> H	87 <sup>g</sup>	
w	H	2-BrC <sub>6</sub> H <sub>4</sub>		(50)
x	H	4-CF <sub>3</sub> C <sub>6</sub> H <sub>4</sub>		(45)

<sup>a</sup> The reactions completed within 24 h except entry d where the conversion was 88%.

<sup>b</sup> The reactions completed within 5 h except in the case of 1 h, which was stirred overnight for completion.

<sup>c</sup> The reactions were performed in acetonitrile with 15 mol % pyrrolidine and 0.3 g/mmol molecular sieves and the data are for the third h.

<sup>d</sup> Addition of 10 mol % of DBU did not result in fulvene formation after 24 h stirring at room temperature.

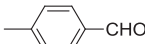
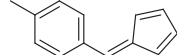

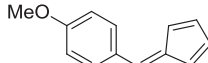
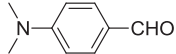
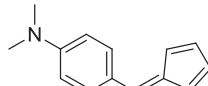
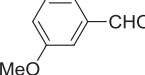
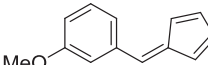
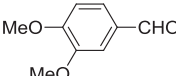
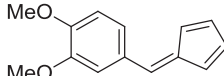
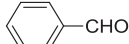
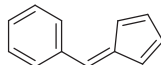
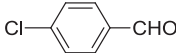
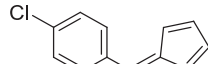

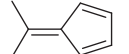
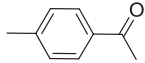
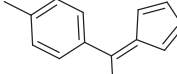

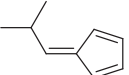
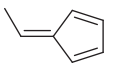

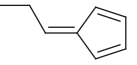

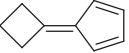

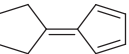

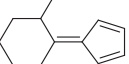


<sup>e</sup> Method B with 3 equiv of added pyrrolidine as catalyst; reaction time 48 h. The yield of the product when no MS was used is 79%.

<sup>f</sup> Reactions not performed; method A: 5 mmol carbonyl compound, 2.5 equiv cyclopentadiene (CP), 1.5 g 3 Å molecular sieves, 10 mol % pyrrolidine, 1.5 equiv NEt<sub>3</sub>, 5 mL MeOH; method B: same as A, no NEt<sub>3</sub>, 1.2 equiv CP.

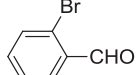
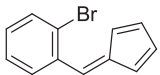

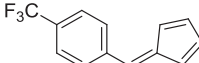
<sup>g</sup> Method A without molecular sieves with 20 mol % of added pyrrolidine as catalyst; isolated yield after overnight stirring and regular aqueous work-up was performed (see Experimental section).

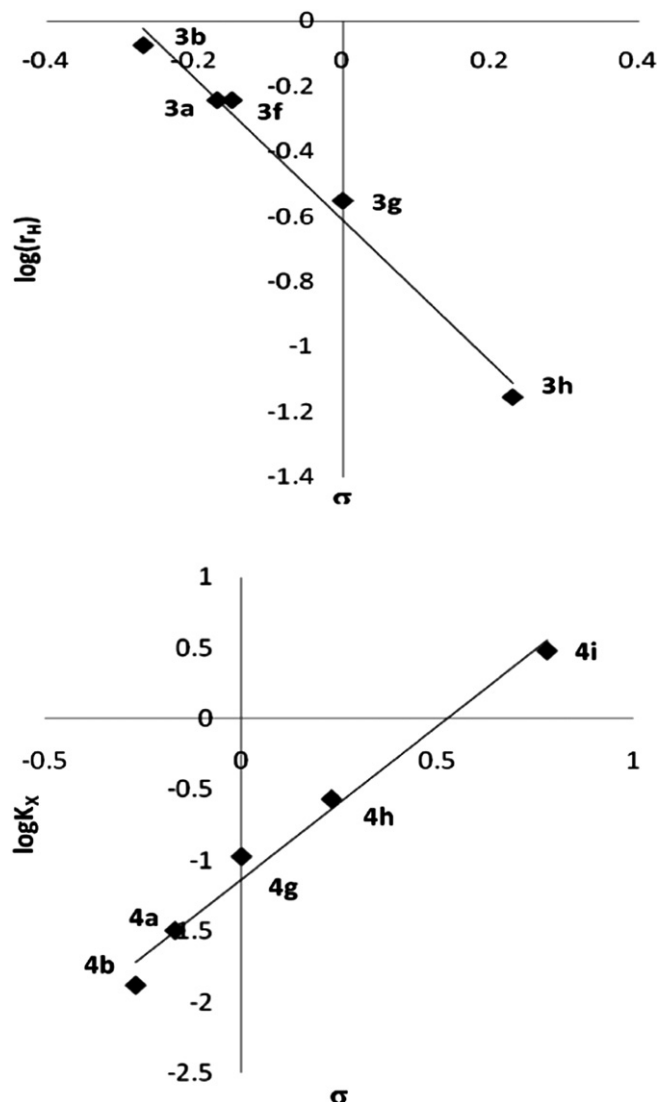
<sup>h</sup> The reaction according to method B gave 10% fulvene. The main product cyclopenta-1,3-dien-1-yl(4-nitrophenyl)methanol and cyclopenta-1,4-dien-1-yl(4-nitrophenyl)methanol (1:1) was isolated in 40% yield after 18 h stirring at room temperature.

**Table 2**  
Method C:<sup>a</sup> optimal conditions (% isolated yields)

Entry	Carbonyl	Fulvene	(%)
a			98
b			97
c			98
e <sup>b</sup>			95
f			98
g <sup>b</sup>			97
h <sup>b</sup>			86
k <sup>b</sup>			77
l <sup>e</sup>			11
n <sup>b</sup>			90
o <sup>b</sup>	CH <sub>3</sub> CHO		90
p <sup>b</sup>			85
q <sup>b</sup>			92
r <sup>b</sup>			95
s <sup>b,c</sup>			44
v <sup>b,d</sup>			60

**Table 2** (continued)

Entry	Carbonyl	Fulvene	(%)
w			25
x			30

<sup>a</sup> Method C: same as B, no molecular sieves, 5 mL MeOH/H<sub>2</sub>O (4:1).<sup>b</sup> To a solution of **1o** (20 mmol) and CP (10.0 mmol), in MeOH/H<sub>2</sub>O (10 mL 7/3) pyrrolidine (10 mol %), was added, and the reaction mixture stirred for 2.5 h at 15 °C. The mixture was poured into an ice-cold mixture of brine solution (10 mL) containing AcOH (1 mmol). The upper oil phase was separated and dried over molecular sieves.<sup>c</sup> To a solution of **2s** (5 mmol) and CP (10 mmol) in 5 mL of methanol/water (4:1), pyrrolidine (10 mmol, 2 equiv) was added, and the mixture stirred at room temperature overnight. The mixture was subjected to flash chromatography on silica gel to give **3s** in 44% yield.<sup>d</sup> Due to anticipated greater water solubility of product, extractive work-up was used.<sup>e</sup> Pyrrolidine (3 equiv) was used: reaction time 24 h.**Fig. 1.** Hammett plot of initial rate  $\log(R_X)$  of fulvene **3a,b,f-h** (top graph) and hemiacetal formation equilibrium constants ( $\log K_X$ ) of **4a,b, g-i** versus  $\sigma$  constants (The correlation equations are  $\log(R_X) = -2.19\sigma - \log(R_X = H)$ ;  $r^2 = 0.99$  and  $\log K_X = 2.16\sigma - \log K_X = H$ ;  $r^2 = 0.98$ , respectively).<sup>26</sup>

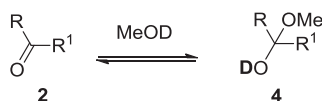
explains the reversed reactivity order in the fulvene formations. The Hammett type plot (Fig. 1) of  $\log K$  versus  $\sigma$  constants produced  $\rho$  constants with the same magnitude as for the fulvene formations but with opposite sign (Scheme 2).

**Table 3**

Substituent effect on the hemiacetal formation equilibria of aromatic aldehydes with MeOD

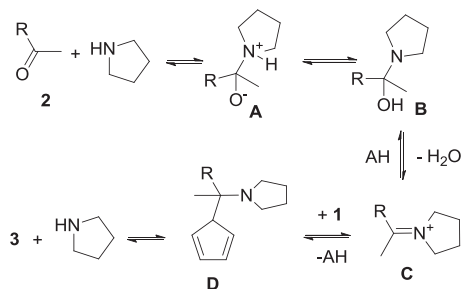
2	$\sigma$	$\log K^a$		
a	-0.17	-1.510	-1.510	-1.495
b	-0.27	-1.440	-1.996	-1.880
g	0	-1.227	-1.061	-0.974
h	0.23	-0.989	-0.630	-0.568
i	0.78	-0.288	0.477	0.477

<sup>a</sup>  $K=[4]/[2]$ ;  $K$  for 1 h, 3.5 h and 24 h.

**Scheme 2.** Hemiactal formation equilibria of aromatic aldehydes.

The correlations discussed above impelled us to take a closer look at the pyrrolidine-catalyzed fulvene formation reaction mechanism.

The Hammett correlations performed confirm the assumption that the rate-determining step of the reaction must involve the formation of a positively charged species. The question was whether it is solely the iminium salt **C** or the C–N bond heterolysis in the transformation of **D** to **3**. The other possibility was that both steps may affect the reaction rate. Our observations with other C–H acids and the aromatic aldehyde series where the addition products are detectable clearly revealed that the addition step of the reaction is faster in the cases of electron-withdrawing substituents. However the substituent effect is reversed for the elimination step of the reaction. The same effect of strong electron-withdrawing substituents was observed under the conditions of methods B and C. The competing formations of cyclopenta-1,3-dien-1-yl(4-aryl)methanol and cyclopenta-1,4-dien-1-yl(4-aryl)methanols **5** were monitored by TLC as well as <sup>1</sup>H NMR. To exclude the possibility of formation of **4** by water addition to the corresponding fulvenes, **3x** was stirred in MeOD in the presence of 20% water and 50 mol % pyrrolidine. No hydration was observed after 24 h. This implies that in the cases of strong electron-withdrawing groups amine displacement competes with amine elimination (Scheme 3). Similarly, the reaction of 4-nitrobenzaldehyde with cyclopentadiene in the presence of only NEt<sub>3</sub> in MeOH-*d*<sub>4</sub> did not give compounds **4i** after 24 h stirring at room temperature. The experiments summarized in Table 4 support a second order reaction with rate= $k[2]^a[1]^b$  where  $a$  and  $b$  are ca. 1.1 and 0.73, respectively. In the present case, the observed substituent effect actually confirms the fact that the rate-limiting step for the overall fulvene formation is the elimination reaction of **D** to give **3** and the catalyst (Table 4).

**Scheme 3.** Equilibria in pyrrolidine-catalyzed condensation of carbonyl compounds with cyclopentadiene.**Table 4**

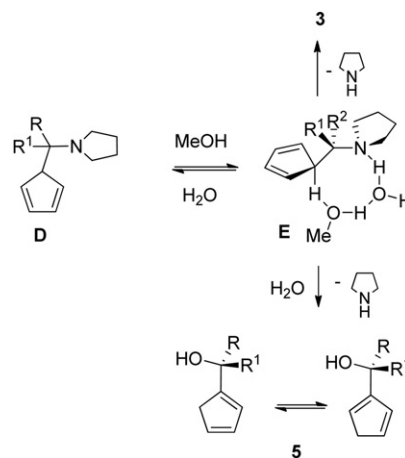
Determination of rate law for the pyrrolidine (5 mol %)-catalyzed fulvene formation between **2b** and **1** in MeOH

<b>2b</b> (mol L <sup>-1</sup> ) <sup>a</sup>	<b>1</b> (mol L <sup>-1</sup> )	Initial rates <sup>b</sup>
1.0	1.0	2.12
1.0	2.0	3.53
2.0	1.0	4.59

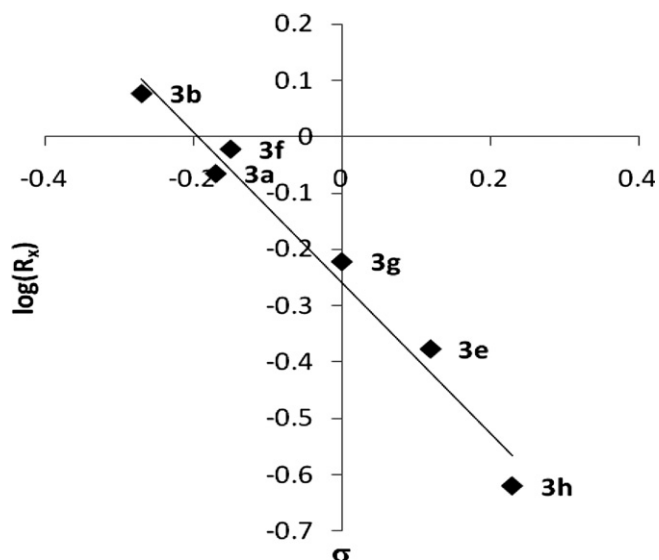
<sup>a</sup> Initial concentration.

<sup>b</sup> The initial rates (mol L<sup>-1</sup> h<sup>-1</sup>) were determined at the 10th min of the reactions at room temperature.

If this conclusion is correct, the reaction without NEt<sub>3</sub> should be much faster because the presence of base may shift the equilibrium from **E** to **C** by deprotonating **E** and/or engaging cyclopentadiene in a hydrogen bond (for solvent effect on the catalyst elimination, see Scheme 4).

**Scheme 4.** Solvent role in catalytic cycle.

To understand the role of NEt<sub>3</sub> on the reaction rate we set up a series of experiments reacting the carbonyl compounds with slight excesses of cyclopentadiene in the presence of only 10 mol % pyrrolidine and molecular sieves (Table 1, method B). To our surprise the absence of NEt<sub>3</sub> enhanced the fulvene formation rates. Again the correlation of the initial rates for the aromatic aldehyde series with Hammett  $\sigma$  constants provided a linear plot with slope ( $\rho$ ) with a lower absolute value (Fig. 2). The comparison of the

**Fig. 2.** Hammett plot of initial rate according to method B ( $\log(R_x)$ ) of fulvene **3a,b, g–i** versus  $\sigma$  constants. The correlation equation is  $\log(R_x) = -1.34\sigma - \log(R_x=H)$   $r^2 = 0.97$ .

reaction susceptibility constants implies that the presence of excess amine decelerates the reaction. This may arise from the direct interaction of intermediate **D** with the amine or an effect on the iminium salt formation. Therefore the use of molecular sieves makes an important difference in the method with a base. In the absence of  $\text{NET}_3$  the rate-limiting step is probably the elimination reaction of **D** to **3**. If so, the use of molecular sieves should not have an effect on the reaction rate (Table 5).

**Table 5**  
Effect of molecular sieves amount and the pyrrolidine concentration on the reaction rate

Entry <sup>a</sup>	MS (g/mmol)	Cat. (mol %)	Solvent	% Conversion <sup>i</sup>		
1	0	10 cat <sup>b</sup>	MeOH	80	100	
2	0.1	10 cat <sup>b</sup>	MeOH	74	100	
3	0.2	10 cat <sup>b</sup>	MeOH	78	100	
4	0.3	10 cat <sup>b</sup>	MeOH	85	100	
5	0.4	10 cat <sup>b</sup>	MeOH	84	100	
6	0.5	10 cat <sup>b</sup>	MeOH	84	100	
7	0.7	10 cat <sup>b</sup>	MeOH	84	100	
8	0	5 cat <sup>b</sup>	MeOH	54	84	100
9	0	1 cat <sup>b</sup>	MeOH	11	28	58
10	0	0.1 cat <sup>b</sup>	MeOH	0	0	0
11	0	0.01 cat <sup>b</sup>	MeOH	0	0	0
12	0	40 cat <sup>c</sup>	MeOH	0.5 <sup>e</sup>		
13	0	80 cat <sup>c</sup>	MeOH	5 <sup>f</sup>		
14	0	10 cat <sup>b</sup>	$\text{CH}_3\text{CN}$	23	55	89 <sup>g</sup>
15	0	10 cat <sup>b</sup>	DMSO	19	48	88 <sup>g</sup>
16	0	10 cat <sup>b</sup>	Ether	0	0	0
17	0	10 cat <sup>b</sup>	THF	0	0	0
18	0	10 cat <sup>b,d</sup>	Ether	0	1	4 <sup>h</sup>
19	0	10 cat <sup>b,d</sup>	THF	0	1	4 <sup>h</sup>

<sup>a</sup> All entries are for the reactions of **2b** with cyclopentadiene in 1:1.2 ratios at room temperature.

<sup>b</sup> Catalyst pyrrolidine.

<sup>c</sup> Catalyst DEA.

<sup>d</sup> MeOH (200 mol %).

<sup>e</sup> The yield after 24 h is 39.

<sup>f</sup> 66% after 24 h.

<sup>g</sup> 100% after 24 h.

<sup>h</sup> 20% after 24 h.

<sup>i</sup> Conversion at 0.5 h, 1.7 h and 5 h, respectively.

Diethylamine was also tested for its catalytic activity (entries 12 and 13) and found that it is less effective than pyrrolidine. However, in spite of the higher concentrations required in some cases, especially in the synthesis of more volatile fulvenes, diethylamine could be preferred due to its lower boiling point.

## 2.2. Solvent effects

Finally, we thought that a study of the solvent effect on the reaction rate would shed some light on the mechanism of the reaction. The solvent series was chosen among those having polarity above and below that of MeOH (entries 1 and 14–17). The reaction rate in acetonitrile was only 1/4th of the rate in MeOH though the polarity of the former is very close to that of MeOH. The rate in much more polar DMSO was again similar to that in acetonitrile. In aprotic solvents as ether and THF no reaction was detected. However, addition of small amounts of MeOH to the latter mixtures provided the formation of **3b** (Table 5, entries 18–19). This clearly proves the catalytic activity of protic solvents in the fulvene formation reactions. Having determined the importance of the solvent to be necessarily protic we have postulated a role for it as depicted in Scheme 4. Intermediate **D** is in equilibrium with **E**, which probably undergoes concerted or near concerted elimination to **3** and the catalyst.<sup>28</sup>

If this postulate is correct, the presence of the more acidic and sterically less demanding water should accelerate the reaction. To prove this we reacted **2a** with cyclopentadiene in water containing

MeOH (Table 6). The result was dramatic: water did not lower the rate. Rather, dramatic rate acceleration was observed at 20% water concentration in MeOH. This clearly proves that the rate-determining step in the reactions without additional base involves the elimination step. However, it seems that this is not the case when sterically encumbered ketones and aldehydes are used (Tables 1 and 2, entries l,w) where the presence of water decelerates the reactions (e.g., the conversions according to method C are 11 and 25% while according to B they are 82 and 50%, respectively, for entries l and w). This is most probably due to the fact that the iminium ion formation equilibrium is being shifted to the left.

**Table 6**  
Effect of water concentration on the reaction rate in the absence of  $\text{NET}_3$

$\text{H}_2\text{O}$ (concn %)	Cat <sup>a</sup> (mol %)	% Conversion (0.25 h)
0	10	35 <sup>b</sup>
2.5	10	35
5	10	37
10	10	39
20	10	62
2.5	200	81

<sup>a</sup> Pyrrolidine.

<sup>b</sup> Product **3a** crystallized in the reaction medium in all cases except the last entry at the 0.5 h and the reactions were completed when checked at 4.5 h by <sup>1</sup>H NMR.

In addition to the accelerating effect discussed above, the use of MeOH/ $\text{H}_2\text{O}$  proved to be of crucial importance in achieving excellent yields in amine-catalyzed synthesis of fulvenes derived from unbranched aliphatic aldehydes by suppressing substrate decomposition due to aldol condensation.

## 2.3. Effect of carbonyl structure and limitations of the amine base promoted fulvene syntheses

The structure of the carbonyl component appears to dictate the success of our amine catalyzed variants of fulvene syntheses. With aldehydes and sterically unencumbered aliphatic ketones pyrrolidine catalysis (method C) appears to be the ideal method. As the steric hindrance increases in the ketone, the use of larger quantities of pyrrolidine (transition to Little–Stone procedure) becomes necessary. As a test substrate for a sterically hindered ketone we used 2-methylcyclohexanone (Table 1, entry s). Whereas methods A and B failed to affect fulvene formation, finally the use of 2 equiv of pyrrolidine (as compared to 5 equiv in Little's procedure) gave the corresponding fulvene in 44% isolated yield. Benzophenone is unreactive under amine base catalysis as well as stoichiometric (or excess base) conditions. However, the use of excess (e.g., 3 equiv) pyrrolidine provided the synthesis of 6-methyl-6-aryl fulvene from the corresponding acetophenone in good yield according to method B and also according to Little's conditions, providing that the reaction times are prolonged (Table 1, entry l). In the cases of diaryl ketones the classical Thiele procedure or a variant thereof with alkoxide base is the method of choice. In the case of keto acids (entries t and u) as well as phenolic substrates (e.g., Table 1, entry d) the best results were obtained with method A (without molecular sieves), still avoiding the use of excess pyrrolidine.

Another important outcome of our studies is the fact that our catalytic method is ideally suited for the synthesis and isolation of fulvenes derived from unbranched aldehydes, in particular those derived from acetaldehyde and propionaldehyde.<sup>29</sup> The stoichiometric (or multiple equivalents) pyrrolidine method using unbranched aldehydes fails to give satisfactory results is presumably due to the fact that excess pyrrolidine in methanol solution results in considerable enamine formation, which induces intermolecular reactions (i.e., aldol condensations) thus rendering the method

unsuitable for these systems. We found that this difficulty can be ameliorated by working in a methanol/H<sub>2</sub>O mixture. It is therefore not surprising to note that particularly 6-alkylfulvenes with short alkyl chains (C1 and 2) that are popular starting materials used in organometallic chemistry are still prepared by the three-step Hafner procedure<sup>7</sup> by way of alkyllithium addition to 6-*N,N*-dimethylaminofulvene instead of the Little procedure. Our catalytic method now offers an excellent direct alternative to Hafner's multistep procedure for the synthesis of the aforementioned 6-alkylfulvenes.

### 3. Conclusions

A systematic investigation of the initial rates of *sec*-amine catalyzed fulvene formations with or without additional NEt<sub>3</sub> and correlations with substituent constants allowed elucidating the mechanism of these reactions. Use of excess secondary amine as well as additional NEt<sub>3</sub> suppressed the reaction rates. In the presence of added NEt<sub>3</sub>, the use of molecular sieves as water capturing agent was shown to be necessary. Our studies have established the following facts: (a) protic solvents, such as MeOH and water accelerate the reaction of cyclopentadiene with carbonyl compounds; (b) the presence of tertiary amines retards reaction rates; (c) electron-donating substituents on the phenyl ring accelerate rates. Based on our mechanistic studies a mechanism involving the interaction of cyclopentadiene with the preformed iminium salt appears reasonable. Based on the studies on the effects of base, MS and solvent we have developed an efficient, catalytic method for the synthesis of a variety of fulvenes. For crystalline products method C is exceptionally simple and practical. The use of MeOH/H<sub>2</sub>O provides the formation of products directly isolable in excellent purity and nearly quantitative yields by vacuum filtration. For liquid fulvenes the procedure is almost as simple. The isolation of the products involves the addition of a specified amount of brine solution containing an equivalent of AcOH at 0 °C resulting in the separation of the organic phase containing practically pure (>95% by <sup>1</sup>H NMR) fulvene that can be pipetted out. Our amine-catalyzed procedure for fulvene synthesis also boasts the added advantage of providing a green alternative to the existing base-promoted methods. Moreover, our alternative methods A and B may be employed in certain cases with success when a phenolic OH or carboxylic acid functionality is present without sacrificing the catalytic use of pyrrolidine. Thus, a wider range of fulvenes, including 6-alkylfulvenes have now become accessible by our catalytic method.

## 4. Experimental section

### 4.1. General

Melting points were determined in open class capillary tubes and are uncorrected. <sup>1</sup>H and <sup>13</sup>C NMR spectra were recorded on 300 MHz and 500 MHz NMR spectrometers, using CDCl<sub>3</sub> as solvent and TMS as internal standard, unless specified otherwise. Most column chromatographic separations were carried out on a flash chromatography system using 40–60 µm silica gel columns using ethyl acetate/*n*-hexane solvent mixtures. For preparative TLC, Merck silica gel (grade 60 PF<sub>254</sub>) was used. All reactions were conducted under an atmosphere of dry nitrogen or argon. Non-deuterated solvents were dried and distilled prior to use. Fresh commercial samples of ethanal (free of trimer) and propanal and 2-methylpropanal were used for fulvene synthesis. Pyrrolidine is an exceptionally foul-smelling and toxic compound and should be handled with care in a well-ventilated hood. Fulvenes are oxygen- and heat-sensitive compounds; all reactions should be carried out under a nitrogen or argon atmosphere. Though all fulvenes

prepared in this study are known compounds, their high resolution NMR spectra are included in the Supplementary data since they are not available in literature in several cases. The initial rates of the reactions used in Hammett<sup>30</sup> correlations were measured at 23 (±1) °C. The change of fulvene or hemiacetal concentrations within the specified periods of time were monitored by <sup>1</sup>H NMR by integrating the appropriate signals of the products and the starting carbonyl compounds.

### 4.2. Synthesis of (2-ylidene)cyclopenta-1,3-dienes 3. Method A. General procedure

To a solution of aldehyde or ketone (5 mmol) and cyclopentadiene (12.5 mmol, 0.8262 g, 1.025 mL) in reagent grade methanol (5 mL), molecular sieve (3 Å, 1.5 g) was added; to this solution pyrrolidine (10 mol %, 0.0355 g) in NEt<sub>3</sub> (7.24 mmol, 0.733 g) was added dropwise within 10 min. The mixture was stirred under nitrogen overnight. The mixture was filtered through sintered glass funnel and the solvent evaporated under vacuum to produce the corresponding fulvene in quantitative yield.

### 4.3. Synthesis of (2-ylidene)cyclopenta-1,3-dienes 3. Method B. General procedure

To a solution of aldehyde or ketone (5 mmol) and cyclopentadiene (6.0 mmol, 0.3966 g, 0.492 mL) in reagent grade methanol (5 mL) molecular sieve (1.5 g, 3 Å) was added, followed by dropwise addition of pyrrolidine (0.0355 g, 10 mol %). The mixture was filtered through sintered glass funnel and the solvent evaporated under vacuum to yield the corresponding fulvene in quantitative yield.

### 4.4. Synthesis of (2-ylidene)cyclopenta-1,3-dienes 3. Method C. General procedure

To a solution of aldehyde or ketone (5 mmol) and cyclopentadiene (0.3966 g, 6.0 mmol/0.492 mL for the crystalline fulvenes or 0.4 mL/5 mmol for the aliphatic substrates) in MeOH/H<sub>2</sub>O (5 mL 4/1) pyrrolidine (0.0355 g, 10 mol %) was added. The formed crystals were filtered, washed with small amounts of cold MeOH/H<sub>2</sub>O mixture and dried in air. For the isolation of oily products the reaction mixture was transferred to an ice-cold mixture of brine solution and 10 mol % AcOH in a narrow graduated cylinder. The organic phase pure was pipetted out. Drying over molecular sieves produced practically pure (>95%) fulvene.

4.4.1. 1-(Cyclopenta-2,4-dienylidenemethyl)-4-methylbenzene (**3a**). (Method C). Orange crystals, mp 60–61 °C. Lit.<sup>31</sup> mp 70.5 °C <sup>1</sup>H NMR (500 MHz, CDCl<sub>3</sub>): δ 7.54 (d, *J*=8.1 Hz, 2H), 7.26 (d, *J*=8.0 Hz, 2H), 7.23 (s, 1H), 6.76 (d, *J*=5.3 Hz, 1H), 6.71–6.68 (m, 1H), 6.55–6.52 (m, 1H), 6.36 (dt, *J*=5.1, 1.7 Hz, 1H), 2.43 (s, 3H). <sup>13</sup>C NMR (126 MHz, CDCl<sub>3</sub>): δ 145.2, 140.2, 139.2, 135.9, 134.8, 131.5, 131.1, 130.2, 128.0, 120.9, 22.1.

4.4.2. 1-(Cyclopenta-2,4-dienylidenemethyl)-4-methoxybenzene (**3b**). (Method C). Orange crystals, mp 68–70 °C. <sup>1</sup>H NMR (500 MHz, CDCl<sub>3</sub>): δ 7.61 (d, *J*=8.6 Hz, 2H), 7.19 (s, 1H), 6.97 (d, *J*=8.8 Hz, 2H), 6.76 (d, *J*=5.3 Hz, 1H), 6.69 (ddd, *J*=5.2, 3.5, 1.6 Hz, 1H), 6.53–6.49 (m, 1H), 6.35 (dt, *J*=5.1, 1.8 Hz, 1H), 3.88 (s, 3H). <sup>13</sup>C NMR (126 MHz, CDCl<sub>3</sub>): δ 161.4, 144.0, 139.0, 135.7, 133.2, 130.52, 130.3, 128.1, 120.6, 115.0, 56.1.

4.4.3. 4-(Cyclopenta-2,4-dienylidenemethyl)-*N,N*-dimethylbenzenamine (**3c**). (Method C). Orange crystals, mp 104–105 °C. Lit.<sup>31</sup> mp 106.5 °C <sup>1</sup>H NMR (500 MHz, CDCl<sub>3</sub>): δ 7.62 (d, *J*=8.7 Hz, 2H), 7.18 (s, 1H), 6.85 (d, *J*=5.2 Hz, 1H), 6.74 (d, *J*=9.0 Hz, 2H), 6.69 (ddd, *J*=5.2,

3.6, 1.6 Hz, 1H), 6.52–6.45 (m, 1H), 6.37 (dt,  $J=5.0, 1.8$  Hz, 1H), 6.37 (dt,  $J=5.0, 1.8$  Hz, 1H), 3.07 (s, 6H).  $^{13}\text{C}$  NMR (126 MHz,  $\text{CDCl}_3$ ):  $\delta$  151.8, 141.5, 140.7, 134.4, 133.6, 128.9, 128.3, 125.4, 120.2, 112.7, 40.8.

4.4.4. 4-(Cyclopenta-2,4-dienylidenemethyl)phenol (**3d**). (Method A). Solid, does not melt below 260 °C  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  7.54 (d,  $J=8.6$  Hz, 2H), 7.16 (s, 1H), 6.88 (d,  $J=8.6$  Hz, 2H), 6.72 (d,  $J=5.2$  Hz, 1H), 6.67 (dd,  $J=5.2, 1.6$  Hz, 1H), 6.50 (d,  $J=4.9$  Hz, 1H), 6.33 (dd,  $J=3.3, 1.7$  Hz, 1H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  156.6, 143.2, 138.1, 134.9, 132.6, 129.8, 129.6, 127.3, 119.8, 115.7. ESI-HRMS calculated for  $\text{C}_{12}\text{H}_{11}\text{O}^+$  ( $\text{M}+\text{H}^+$ ) 171.0804, found 171.0895.

4.4.5. 1-(Cyclopenta-2,4-dienylidenemethyl)-3-methoxybenzene (**3e**). (Method C).  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  7.34 (t,  $J=7.9$  Hz, 1H), 7.19 (d,  $J=8.7$  Hz, 2H), 7.13 (s, 1H), 6.93 (dd,  $J=8.2, 2.1$  Hz, 1H), 6.69 (dt,  $J=5.3, 3.5$  Hz, 2H), 6.53 (d,  $J=5.0$  Hz, 1H), 6.34 (dd,  $J=3.4, 1.7$  Hz, 1H), 3.86 (s, 3H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  159.6, 145.3, 138.0, 135.4, 130.9, 129.6, 127.1, 123.2, 120.3, 115.56, 114.9, 55.2.

4.4.6. 4-(Cyclopenta-2,4-dienylidenemethyl)-1,2-dimethoxybenzene (**3f**). (Method C). Orange crystals, mp 72–73 °C. Lit.<sup>32</sup> mp 72 °C  $^1\text{H}$  NMR (500 MHz,  $\text{CDCl}_3$ ):  $\delta$  7.23–7.19 (m, 2H), 7.18 (s, 1H), 6.93 (d,  $J=8.2$  Hz, 1H), 6.75 (d,  $J=5.3$  Hz, 1H), 6.70 (ddd,  $J=5.2, 3.5, 1.6$  Hz, 1H), 6.51 (dd,  $J=3.7, 1.4$  Hz, 1H), 6.35 (dt,  $J=5.1, 1.8$  Hz, 1H), 3.95 (s, 3H), 3.95 (s, 3H).  $^{13}\text{C}$  NMR (126 MHz,  $\text{CDCl}_3$ ):  $\delta$  151.1, 149.8, 144.2, 139.2, 135.8, 130.6, 130.5, 128.2, 125.5, 120.5, 113.9, 111.9, 56.7, 56.6.

4.4.7. 1-(Cyclopenta-2,4-dienylidenemethyl)benzene (**3g**). (Method C).  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  7.61 (d,  $J=6.9$  Hz, 2H), 7.51–7.32 (m, 3H), 7.24 (s, 1H), 6.76–6.64 (m, 2H), 6.54 (d,  $J=5.1$  Hz, 1H), 6.35 (dd,  $J=3.4, 1.7$  Hz, 1H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  145.2, 138.1, 136.7, 135.4, 130.8, 130.6, 128.9, 128.6, 127.1, 120.3.

4.4.8. 1-Chloro-4-(cyclopenta-2,4-dienylidenemethyl)benzene (**3h**). (Method C). Red crystals, mp 43–45 °C.  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  7.52 (d,  $J=8.5$  Hz, 2H), 7.40 (d,  $J=8.5$  Hz, 2H), 7.16 (s, 1H), 6.68 (dt,  $J=13.4, 3.3$  Hz, 2H), 6.54 (d,  $J=4.8$  Hz, 1H), 6.33 (dd,  $J=3.6, 1.6$  Hz, 1H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  145.6, 136.4, 135.8, 135.1, 135.1, 131.7, 131.1, 128.9, 127.1, 119.9.

4.4.9. 1-(Cyclopenta-2,4-dien-1-ylidenemethyl)-4-nitrobenzene (**3i**). To a solution of 4-nitrobenzaldehyde (5 mmol) and cyclopentadiene (6.0 mmol, 0.3966 g, 0.492 mL) in reagent grade acetonitrile (5 mL) molecular sieve (1.5 g, 3 Å) was added, followed by dropwise addition of pyrrolidine (0.0532 g, 15 mol %). The mixture was stirred overnight and filtered through sintered glass funnel and the solvent evaporated under vacuum. The mixture was separated by flash chromatography using silica packed column and hexane/ethyl acetate as eluent. For further purification preparative TLC was used. Reddish oil.  $^1\text{H}$  NMR (500 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.27 (1H, d,  $J=9$  Hz), 7.70 (1H, d,  $J=9$  Hz), 7.20 (1H, s), 6.81–6.63 (1H, m), 6.57 (1H, dd,  $J=5, 2$  Hz), 6.38–6.26 (1H, m).  $^{13}\text{C}$  NMR (126 MHz,  $\text{CDCl}_3$ ):  $\delta$  148.6, 147.8, 143.3, 137.5, 134.6, 132.9, 131.2, 127.2, 124.1, 120.1.

4.4.10. 1-(Cyclopenta-2,4-dien-1-ylidenemethyl)-3-nitrobenzene (**3j**). Prepared and purified as **3i**. Reddish oil.  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.42 (1H, s), 8.21 (1H, dd,  $J=8, 1$  Hz), 7.86 (1H, d,  $J=8$  Hz), 7.60 (1H, t,  $J=8$  Hz), 7.20 (1H, s), 6.76–6.67 (1H, m), 6.61 (1H, d,  $J=5$  Hz), 6.56 (1H, d,  $J=5$  Hz), 6.36–6.28 (1H, m).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  148.3, 147.4, 138.2, 136.9, 135.9, 134.2, 132.2, 129.5, 126.9, 124.8, 123.2, 119.6.

4.4.11. 5-(Propan-2-ylidene)cyclopenta-1,3-diene (**3k**). (Method C).  $^1\text{H}$  NMR (500 MHz,  $\text{CDCl}_3$ ):  $\delta$  6.58–6.56 (m, 2H), 6.52 (dd,  $J=4.0,$

2.3 Hz, 2H), 2.25 (s, 6H).  $^{13}\text{C}$  NMR (126 MHz,  $\text{CDCl}_3$ ):  $\delta$  150.8, 143.3, 131.4, 121.3, 23.8.

4.4.12. 1-(1-(Cyclopenta-2,4-dien-1-ylidene)ethyl)-4-methylbenzene (**3l**). Prepared according to method B by using 3 equiv of pyrrolidine. Reddish oil.  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  7.31 (2H, d,  $J=8$  Hz), 7.21 (2H, d,  $J=8$  Hz), 6.67–6.62 (1H, m), 6.57 (1H, d,  $J=5$  Hz), 6.49 (1H, d,  $J=5$  Hz), 6.24–6.20 (1H, m), 2.54 (3H, s), 2.40 (3H, s).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  150.0, 143.0, 139.0, 138.2, 131.5, 131.2, 129.2, 128.5, 123.7, 121.0, 22.5, 21.2.

4.4.13. 5-(2-Methylpropylidene)cyclopenta-1,3-diene (**3m**). (Method C). Yellow oil.  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  6.54 (t,  $J=3.6$  Hz, 2H), 6.48 (d,  $J=5.0$  Hz, 1H), 6.26 (d,  $J=10.1$  Hz, 1H), 6.23–6.19 (m, 1H), 3.11–2.94 (m, 1H), 1.15 (d,  $J=6.6$  Hz, 6H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  149.7, 143.4, 132.8, 130.6, 125.8, 119.1, 30.3, 23.0.

4.4.14. 5-Ethylidene-cyclopenta-1,3-diene (**3o**). (Method C). Yellow oil.  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  6.61 (t,  $J=1.6$  Hz, 2H), 6.58–6.48 (m, 2H), 6.26 (dt,  $J=5.1, 1.7$  Hz, 1H), 2.21 (d,  $J=7.3$  Hz, 3H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  146.9, 137.8, 132.9, 130.6, 125.4, 118.8, 16.7.

4.4.15. 5-Propylidene-cyclopenta-1,3-diene (**3p**). (Method C). Yellow oil. Yield 85%.  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  6.55 (dd,  $J=4.7, 3.5$  Hz, 2H), 6.45 (dd,  $J=18.2, 6.5$  Hz, 2H), 6.23 (dd,  $J=3.5, 1.6$  Hz, 1H), 2.57 (p,  $J=7.6$  Hz, 2H), 1.17 (t,  $J=7.5$  Hz, 3H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  145.2, 144.6, 132.9, 130.7, 125.6, 118.9, 24.3, 14.0.

4.4.16. 5-Cyclobutylidene-cyclopenta-1,3-diene (**3q**). (Method C). Yellow oil.  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  6.45 (d,  $J=5.2$  Hz, 2H), 6.31 (dd,  $J=4.4, 2.0$  Hz, 2H), 3.14 (t,  $J=7.8$  Hz, 4H), 2.18 (p,  $J=7.9$  Hz, 2H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  160.3, 138.7, 130.2, 119.8, 32.2, 16.9.

4.4.17. 5-Cyclopentylidene-cyclopenta-1,3-diene (**3r**). (Method C). Yellow oil.  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  6.44 (dt,  $J=6.2, 2.9$  Hz, 4H), 2.84 (t,  $J=7.2$  Hz, 4H), 1.91–1.73 (m, 4H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  162.5, 138.3, 130.0, 121.3, 33.0, 26.0.

4.4.18. 1-(Cyclopenta-2,4-dien-1-ylidene)-2-methylcyclohexane (**3s**). (Method C). Yellow oil.  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  6.70–6.45 (m, 4H), 3.48–3.32 (m, 1H), 2.94 (d,  $J=14.0$  Hz, 1H), 2.49 (td,  $J=13.7, 4.9$  Hz, 1H), 2.06–1.93 (m, 1H), 1.86–1.43 (m, 5H), 1.29 (d,  $J=7.2$  Hz, 3H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  162.4, 139.2, 130.6, 120.1, 119.5, 35.1, 33.7, 28.9, 28.4, 20.5, 19.5.

4.4.19. 4-(Cyclopenta-2,4-dien-1-ylidene)pentanoic acid (**3t**). (Method A without MS). Yellow solid mp 65–67 °C. Lit.<sup>12</sup> mp 66–67 °C.  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  10.9 (br s, 1H), 6.52 (d,  $J=5.7$  Hz, 4H), 2.95–2.82 (m, 2H), 2.62 (dd,  $J=8.8, 7.0$  Hz, 2H), 2.22 (s, 3H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  178.9, 149.6, 143.4, 131.5, 131.3, 120.7, 120.1, 33.4, 31.4, 20.5.

4.4.20. 5-(Cyclopenta-2,4-dien-1-ylidene)hexanoic acid (**3u**). (Method A without MS). Yellow oil, which solidifies on standing in the fridge.  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  10.01 (br s, 1H), 6.58–6.41 (m, 4H), 2.60 (t,  $J=7.7$  Hz, 2H), 2.41 (t,  $J=7.4$  Hz, 2H), 2.22 (s, 3H), 1.93 (p,  $J=7.5$  Hz, 2H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ ):  $\delta$  179.5, 151.6, 143.2, 131.1, 131.0, 120.6, 120.2, 35.7, 33.35, 23.6, 20.6. ESI-HRMS calculated for  $\text{C}_{11}\text{H}_{15}\text{O}_2^+$  ( $\text{M}+\text{H}^+$ ) 179.1067, found 179.1086.

4.4.21. 2-(Cyclopenta-2,4-dienylidene)propan-1-ol (**3v**). (Method C). Yellow oil, which slowly solidifies in the fridge. Mp 52–53 °C. Lit.<sup>33</sup> mp 54–54.5 °C  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  6.52 (dd,  $J=6.2, 4.0$  Hz, 4H), 4.54 (s, 2H), 2.28 (s, 3H), 1.91 (br s, 1H).  $^{13}\text{C}$  NMR

(75 MHz, CDCl<sub>3</sub>):  $\delta$  149.0, 142.5, 132.3, 131.7, 121.2, 119.5, 64.6, 18.3.

4.4.22. *1-Bromo-2-(cyclopenta-2,4-dien-1-ylidenemethyl)benzene (3w)*. Reddish oil. <sup>1</sup>H NMR (300 MHz, CDCl<sub>3</sub>):  $\delta$  7.62 (2H, ddd,  $J=14, 8, 1$  Hz), 7.41 (1H, s), 7.38–7.30 (1H, m), 7.22 (1H, td,  $J=8, 2$  Hz), 6.67 (1H, ddd,  $J=5, 3, 2$  Hz), 6.57 (1H, d,  $J=5$  Hz), 6.49 (1H, d,  $J=5$  Hz), 6.43–6.36 (1H, m) <sup>13</sup>C NMR (75 MHz, CDCl<sub>3</sub>):  $\delta$  146.4, 136.4, 135.5, 132.9, 132.8, 131.8, 129.9, 127.2, 126.4, 120.5. ESI-HRMS calculated for C<sub>12</sub>H<sub>10</sub>Br<sup>+</sup> (M+H<sup>+</sup>) 232.9960, not observed.

4.4.23. *1-(Cyclopenta-2,4-dien-1-ylidenemethyl)-4-(trifluoromethyl)benzene (3x)*. Red crystals mp 60–62 °C. <sup>1</sup>H NMR (300 MHz, CDCl<sub>3</sub>):  $\delta$  7.67 (4H, s), 7.21 (1H, s), 6.73–6.66 (1H, m), 6.61 (1H, d,  $J=5$  Hz), 6.56 (1H, d,  $J=5$  Hz), 6.33 (1H, dd,  $J=4, 2$  Hz) <sup>13</sup>C NMR (75 MHz, CDCl<sub>3</sub>):  $\delta$  147.0, 140.0, 136.4, 135.6, 131.9, 130.5, 126.9, 125.42 (q,  $J=3.75$  Hz), 120.0.

4.4.24. *Cyclopenta-1,3-dien-1-yl(4-nitrophenyl)methanol compound with cyclopenta-1,4-dien-1-yl(4-nitrophenyl)methanol (1:1) (4i)*. Yellowish oil. <sup>1</sup>H NMR (300 MHz, CDCl<sub>3</sub>):  $\delta$  8.15 (2H, d,  $J=9$  Hz), 7.54 (2H, dd,  $J=8, 7$  Hz), 6.49–6.20 (3H, m), 5.73 (1H, d,  $J=12$  Hz), 3.01 (1H, t,  $J=3$  Hz), 3.11–2.59 (2H, m) <sup>13</sup>C NMR (75 MHz, CDCl<sub>3</sub>):  $\delta$  150.6, 149.9, 149.0, 148.0, 147.0, 147.0, 135.3, 133.8, 131.6, 130.9, 130.0, 128.9, 127.0, 126.8, 123.5, 72.0, 71.4, 41.3, 40.0. ESI-HRMS calculated for C<sub>12</sub>H<sub>10</sub>Br<sup>+</sup> (M+H<sup>+</sup>) 232.9960, not observed.

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## Supplementary data

<sup>1</sup>H and <sup>13</sup>C NMR spectra of compounds **3a–h**, **k**, **n–s** and **v**. Supplementary data related to this article can be found online at doi:10.1016/j.tet.2011.09.036.

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